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Forms of Energy Answers

Fill in the blanks with the words at the bottom of the page. Some words will be used more than once. Use the word that best completes the sentence.

- Stored energy and the energy of position are potential energy.
- Compressed springs and stretched rubber bands are examples of elastic energy.
- The vibration and movement of the atoms and molecules within substances is called thermal energy.
- The scientific rule that states that energy cannot be created or destroyed is called the Law of Conservation of Energy.
- The movement of energy through substances in longitudinal waves is sound energy.
- The energy of position—such as a rock on a hill—is gravitational potential energy.
- The movement of objects and substances from place to place is motion energy.
- Electromagnetic energy traveling in transverse waves is radiant energy.
- Energy stored in the bonds of atoms and molecules is chemical energy.
- The movement of atoms, molecules, waves, and electrons is kinetic energy.
- The movement of electrons is electrical energy.
- The amount of useful energy you get from a system is its energy efficiency.
- The energy in petroleum and coal is stored as chemical energy.
- X-rays are an example of radiant energy.
- Fission and fusion are examples of nuclear energy.
- A hydropower reservoir is an example of gravitational potential energy.
- Wind is an example of the energy of motion.

Word Bank

*chemical	*electrical	*kinetic	*potential	*thermal
*Conservation of Energy	*energy efficiency	*motion	*radiant	
*elastic	*gravitational potential	*sound		

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ANSWERS

1. What was the first machine that could add, subtract, multiply, and divide?
A. the abacus
B. the mechanical calculator
C. the arithmometer
D. the modern computer
2. What component of the computer is correctly associated with the function it performs?
A. monitor: output
B. microprocessor : storage
C. random access memory (RAM): processing
D. optical drive: input
3. Which one of these is not an input device?
A. speakers
B. mouse
C. scanner
D. keyboard
4. What type of storage does not accept new information?
A. hard drive
B. removable media
C. random access memory (RAM)
D. read only memory (ROM)
5. Microprocessors are used in:
A. DVD players
B. computers
C. calculators
D. remote controls
6. What is not a way computers can output information?
A. printer
B. monitor
C. mp3 files

Katherine Bailey

Reading Assignment 1
Multiple Choice Test

EDN303 J. Robertson

Introduction to computers

Lesson 1

1. What was the first machine that could add, subtract, multiply, and divide?
A. the abacus
B. the mechanical calculator
C. the arithmometer
D. the modern computer

Lesson 2

2. What component of the computer is correctly associated with the function it performs?
A. monitor: output
B. microprocessor : storage
C. random access memory (RAM): processing
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Lesson 3

3. Which one of these is not an input device?
A. speakers
B. mouse
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Lesson 4

4. What type of storage does not accept new information?
A. hard drive
B. removable media
C. random access memory (RAM)
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Lesson 5

5. Microprocessors are used in:
A. DVD players
B. computers
C. calculators
D. remote controls

Lesson 6

6. What is not a way computers can output information?
A. printer
B. monitor
C. mp3 files

COUNTING ATOMS

ANSWER KEY

1. How many atoms are in each molecule?
a. H₂O: 3 atoms (2 H, 1 O)
b. CO₂: 3 atoms (1 C, 2 O)
c. CH₄: 5 atoms (1 C, 4 H)
d. H₂S: 3 atoms (2 H, 1 S)
e. HF: 2 atoms (1 H, 1 F)
f. NH₃: 4 atoms (1 N, 3 H)

2. How many molecules are in each sample?
a. 10 H₂O molecules
b. 5 CO₂ molecules
c. 2 CH₄ molecules
d. 3 H₂S molecules
e. 1 HF molecule
f. 1 NH₃ molecule

Covalent bonding

A non-metal and another non-metal can bond together by sharing electrons so that both electrons can achieve a full outer shell. This is called covalent bonding. To represent the atoms involved in the bonding we draw circles for the electrons on one of the atoms and dots for the electrons on the other atoms as shown here.

Instructions

NB: It may be easiest to draw each atom out individually first to work out how many electrons will be shared.

1. Draw the circles (outer shells) and overlap them (like a Venn diagram).
2. Write the symbols of the atoms involved in the middle of each circle.
3. Draw the electrons for one atom as dots (•) and electrons for the other atom as crosses (x).
4. Any shared electrons are drawn in the overlapping section.
5. Count up the electrons for EACH atom and check that they each now have a full outer shell (2 electrons for the first shell and 8 for the next 2 shells).

TABLE: Draw dot and cross diagrams for the following covalent compounds.

F ₂	CH ₄
HCl	CO ₂
H ₂ S	H ₂ O
HF	NH ₃

Challenge - O₂

Our expert science tutors break down the topics through interactive one-to-one sessions. The first one-on-one online tutoring lesson is always FREE, no purchase obligation, no credit card required. For example, two atoms of hydrogen combine with one atom of oxygen to form a molecule named water. Compounds may be solids, liquids, or gases. Water is thus a compound – it always has two atoms of hydrogen and one atom of oxygen in each molecule, no more no less. You may schedule online tutoring lessons at your personal scheduled times, all with a Money-Back Guarantee. Table salt contains an equal number of positively charged sodium ions and negatively charged chloride ions. For answers/solutions to any question or to learn concepts, take a FREE TRIAL Session. Organic Compounds are generally made up of carbon and hydrogen. Click on the links in the rest of the article to discover worksheets for each of the topics addressed. This simple display of subtraction will help guide students through the process of early arithmetic, which can be additionally aided by these subtraction facts to 10. Table salt, formed from sodium and chlorine, is an ionic compound. All other compounds that lack Carbon-Hydrogen bonds are generally called Inorganic Compounds. When working with first-grade students, it's important to start from where they understand and work your way up, ensuring that each student masters each concept individually before moving on to the next topic. eTutorWorld offers affordable one-on-one live tutoring over the web for Grades K-12, Test Prep help for Standardized tests like SCAT, CogAT, MAP, SSAT, SAT, ACT, ISEE and AP. Atoms of the same elements combine to form molecules while atoms of different elements combine to form compounds. A molecule is the smallest particle that retains all the properties of the substance from which it is obtained. An ion with a positive charge, as sodium in the example above, is called a Cation while an ion with a negative charge, as chloride in the example above, is called an Anion. First-grade teachers may also introduce their students to a base-level knowledge of fractions, geometric shapes, and mathematical patterns, though none of them are required course material until the second and third grades. Additionally, students will be expected to recognize number patterns and should practice their skills in counting by 2s, counting by 5s, and counting by 10s and identifying whether a number is greater than or less than 20, and be able to parse out mathematical equations from word problems like these, which may include ordinal numbers up to 10 In terms of practical math skills, the first grade is also an important time to ensure students understand how to tell time on a clock face and how to count U.S. coins up to 50 cents. For example, two oxygen atoms combine to form an oxygen molecule. These removed electrons may then join other atoms or groups of atoms to make them Negative Ions. In ionization process most metals lose electrons to form positive ions while most non-metals gain electrons to form negative ions. If the number of electrons and protons are equal, the atom is neutral. Halogens such as Fluorine, Chlorine, and Iodine, always form anions Alkali metals such as Lithium, Sodium, Potassium, always form cations Alkali earth metals such as Barium, Calcium, Magnesium, always form cations Noble gases are generally unreactive and tend not to form ions Most other metals such as Iron, Nickel, Silver, form cations, while most non-metals, such as Oxygen, Sulfur, Carbon, form anions. An atom with an electric charge is called an ion. No credit card required, no obligation to purchase. They cannot be broken down further by chemical means. (A negatively charged chlorine atom is called a Chloride). Compounds hold their atoms together in many ways, such as covalent compounds like water and methane are held together by covalent bonds Ionic compounds like sodium chloride (table salt) and sodium fluoride (found in toothpaste) are held together by ionic bonds Metallic compounds like gold and aluminum are held together by metallic bonds Compounds have an identifiable structure based on how their atoms are arranged. Every atom has negatively charged electrons spinning around a heavy nucleus with positively charged protons. Water molecule with an oxygen atom (shown red) and two hydrogen atoms (shown blue) Learn more about Structure of Matter and other important topics with 8th Grade Science Tutoring at eTutorWorld. Sodium Chloride has a fixed crystalline structure consisting of sodium cations (shown in purple) and chloride anions (shown in green). Atoms of the same element may combine to form Molecules. Each compound has its own unique properties, which are different from the properties of the individual elements from which it is made. The chemical process of removing electrons from atoms or molecules to produce Positive Ions is called Ionization. Explore more concepts in these extra worksheets: Further division of the molecule will yield only atoms of individual elements. Assigning worksheets like "Order the Numbers to 50" will help teachers assess whether or not a student fully grasps the number line. Just schedule a FREE Session to meet a tutor and get help on any topic you want! When atoms of different elements combine in a fixed ratio, the result is named Compound. First-grade math students will be introduced to basic addition and subtraction, oftentimes in the form of word problems, over the

course of the year, meaning they will be expected to add up to 20 and subtract numbers below fifteen, both of which won't require the students to re-group or "carry the one." These concepts are easiest understood through tactile demonstration such as number blocks or tiles or through illustration or example such as showing the class a pile of 15 bananas and taking away four of them, then asking the students to calculate then count the remaining bananas. For example, carbon forms a range of Organic Compounds with hydrogen, such as Methane, Ethane, Propane, Butane, to name a few, each with its unique structure of atom arrangement. One of the first things first graders have to master is the concept of counting to 20, which will help them quickly count beyond those basic numbers and begin to understand the 100s and 1000s by the time they reach the second grade. Atoms of two different elements may also combine to form a molecule. The diagrams show carbon atoms in green and hydrogen in grey. Just schedule a FREE Sessions to meet a tutor and get help on any topic you want! An atom or a group of atoms (a molecule) with an electric charge is called an ion. Just schedule a FREE Sessions to meet a tutor and get help on any topic you want! Tutoring Package Validity Grade (1-12), College 5 sessions 1 Month \$124 1 session 1 Month \$25 10 sessions 3 months \$239 15 sessions 3 months \$354 20 sessions 4 months \$449 50 sessions 6 months \$1049 100 sessions 12 months \$2049 Images Credit: ©2022 eTutorWorld Terms of use Privacy Policy Site by Little Red Bird ©2022 eTutorWorldTerms of usePrivacy PolicySite by Little Red Bird When it comes to teaching first-grade students the common core standards of mathematics, there's no better way to practice than with worksheets geared toward repeatedly applying the same basic concepts such as counting, adding and subtracting without carrying, word problems, telling time, and calculating currency. No credit card required, no obligation to purchase. The amount of charge that an ion has is determined by the number of electrons gained or lost by the atom or molecule. These skills will be essential as students begin to apply two-digit addition and subtraction in the second grade. Each sodium atom loses an electron to become a positively charged sodium ion and each chlorine atom gains an electron to become a negatively charged chloride ion. Check out "Understanding 1/2," this "Shape Book," and these additional 10 Geometry worksheets for late Kindergarten and Grade 1. When working with first-grade students, it's important to start from where they are. We also offer the advantage of customized lesson plans, flexible schedules and convenience of learning from home. Here's another way to ask the question: A man was holding some balloons and the wind blew 4 away, How many are left? However, in some cases, students may require additional attention or explanation beyond what worksheets alone can offer—for this reason, teachers should also prepare demonstrations in class to help guide students through the coursework. It is also important to focus on thinking concepts. _____ is a negatively charged ion while _____ is a positively charged ion. Molecules, Compounds Carbon, Hydrogen Ion Anion, Cation Cations, Anions No credit card required, no obligation to purchase. Too often we ask questions where the unknown is at the end of the question, but the unknown can also be put at the beginning of the question. As young mathematicians progress through their early education, they will be expected to demonstrate comprehension of these basic skills, so it's important for teachers to be able to gauge their students' aptitudes in the subject by administering quizzes, working one on one with each student, and by sending them home with worksheets like the ones below to practice on their own or with their parent. He only has 6 balloons left, how many did he start with? The Periodic Table arranges all known elements in the order of their Atomic Numbers. For instance, think about this word problem: A man has 10 balloons and the wind blew 4 away. Students will also be expected to demonstrate a comprehension of addition, through completing word problems that feature addition sentences up to 10, and worksheets like " Adding to 10," " Adding to 15, " and "Adding to 20" will help teachers gauge students' comprehension of the basics of simple addition. An element is a pure substance consisting of atoms which have the same number of protons or atomic number. Elements are the most basic chemical substances.

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